

1. Which statement best explains why nitrogen has a larger first ionisation energy than oxygen?
- A N atoms have less repulsion between p-orbital electrons than O atoms.
 - B N atoms have a smaller nuclear charge than O atoms.
 - C N atoms lose an electron from the 2s subshell, while O atoms lose an electron from the 2p subshell.
 - D N atoms have an odd number of electrons, while O atoms have an even number.

Your answer

[1]

END OF QUESTION PAPER